

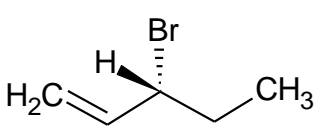
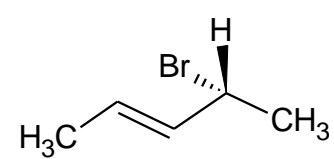
Class XII
Chemistry (Code – 043)
Sample Question Paper 2018-19

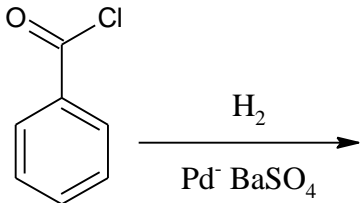
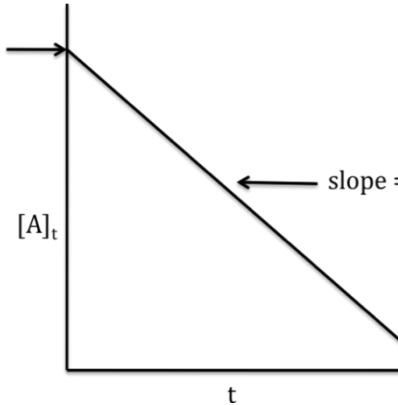
Time allowed: 3 Hours

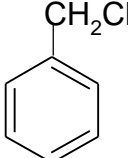
Max. Marks: 70

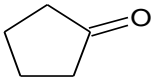
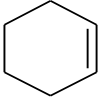
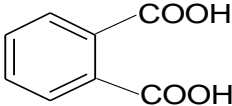
General Instructions:


- (a) All questions are compulsory.
- (b) Section A: Q.no. 1 to 5 are very short answer questions and carry 1 mark each.
- (c) Section B: Q.no. 6 to 12 are short answer questions and carry 2 marks each.
- (d) Section C: Q.no. 13 to 24 are also short answer questions and carry 3 marks each.
- (e) Section D: Q.no. 25 to 27 are long answer questions and carry 5 marks each.
- (f) There is no overall choice. However an internal choice has been provided in two questions of one mark, two questions of two marks, four questions of three marks and all the three questions of five marks weightage. You have to attempt only one of the choices in such questions.
- (g) Use of log tables if necessary, use of calculators is not allowed.

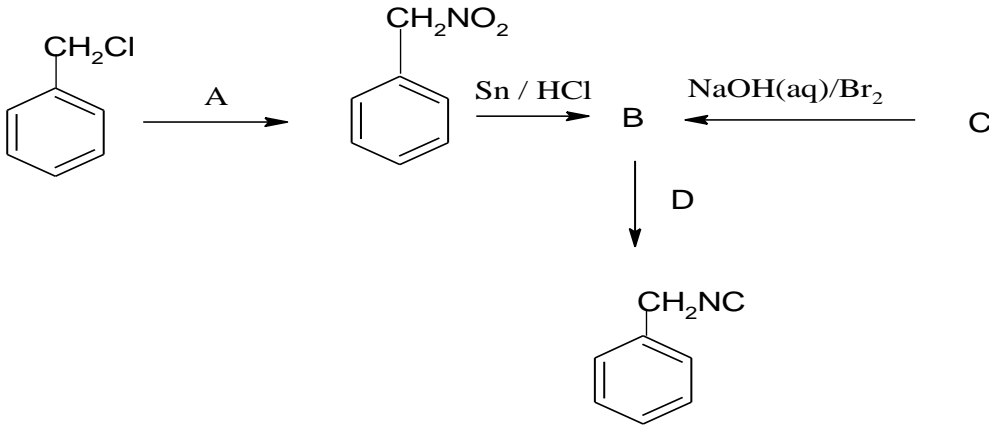
Section-A		
1.	ZnO crystal on heating acquires the formula Zn_{1-x}O . Give reason.	1
	OR	
	There is an increase in conductivity when Silicon is doped with Phosphorous. Give reason	
2.	Based on the type of dispersed phase, what type of colloids are micelles?	1
3.	On the basis of crystal field theory, write the electronic configuration of d^6 in terms of t_{2g} and e_g in an octahedral field when $\Delta_o < P$.	1
	OR	
	Low spin configuration are rarely observed in tetrahedral coordination entity formation. Explain	
4.	Identify the compound that on hydrogenation produces an optically active compound from the following compounds: <div style="display: flex; justify-content: space-around; align-items: center;"> <div style="text-align: center;">  <p>(A)</p> </div> <div style="text-align: center;">  <p>(B)</p> </div> </div>	1
5.	Write the name of the biodegradable polymer used in orthopaedic devices.	1

Section-B		
6.	Calculate the freezing point of a solution containing 8.1 g of HBr in 100 g of water, assuming the acid to be 90 % ionized. [Given: Molar mass Br = 80 g/mol, K_f water = 1.86 K kg / mol]	2
	OR	
	Calculate the molality of ethanol solution in which the mole fraction of water is 0.88.	
7.	Identify the reaction and write the IUPAC name of the product formed: (a)	2
	$\text{CH}_3\text{-CH}_2\text{-COOH} \xrightarrow{\text{(i) Br}_2 / \text{Red phosphorous}}$	
	(b) 	
8.	OR	2
	Write the structures and IUPAC names of the cross aldol condensation products only of ethanal and propanal.	
9.	(a) Justify the role of tert-butyl peroxide in the polymerization of ethene. (b) Write the structures of the monomers of the following polymer:	2
	$\left[\text{CH}_2 - \text{CH} = \text{CH} - \text{CH}_2 - \text{CH}_2 - \overset{\text{C}_6\text{H}_5}{\underset{ }{\text{CH}}} \right]_n$	
10.	Write the mechanism of acid dehydration of ethanol to yield ethane	2
10.	For a certain chemical reaction variation in concentration [A] vs. time (s) plot is given below:	
	(i) Predict the order of the given reaction?	
	(ii) What does the slope of the line and intercept indicate?	
	(iii) What is the unit of rate constant k?	
		

11.	<p>Draw the molecular structures of the following:</p> <p>(a) Noble gas species which is isostructural with BrO_3^-</p> <p>(b) Dibasic oxoacid of phosphorus</p>	2
12.	<p>(i) On the basis of the standard electrode potential values stated for acid solutions, predict whether Ti^{4+} species may be used to oxidise Fe(II) to Fe(III)</p> $Ti^{4+} + e^- \rightarrow Ti^{3+} \quad E^o = +0.01V$ $Fe^{3+} + e^- \rightarrow Fe^{2+} \quad E^o = +0.77V$ <p>(ii) Based on the data arrange Fe^{2+}, Mn^{2+} and Cr^{2+} in the increasing order of stability of +2 oxidation state. (Give a brief reason)</p> $E^o_{Cr^{3+}/Cr^{2+}} = -0.4V$ $E^o_{Mn^{3+}/Mn^{2+}} = +1.5V$ $E^o_{Fe^{3+}/Fe^{2+}} = +0.8V$	2
Section-C		
13.	Niobium crystallises in body-centred cubic structure. If the atomic radius is 143.1 pm, calculate the density of Niobium. (Atomic mass = 93u).	3
14.	<p>Give reasons for the following:</p> <p>a. When 2g of benzoic acid is dissolved in 25 g of benzene, the experimentally determined molar mass is always greater than the true value.</p> <p>b. Mixture of ethanol and acetone shows positive deviation from Raoult's Law.</p> <p>c. The preservation of fruits by adding concentrated sugar solution protects against bacterial action.</p>	3
15.	An alcohol A ($C_4H_{10}O$) on oxidation with acidified potassium dichromate gives acid B ($C_4H_8O_2$). Compound A when dehydrated with conc. H_2SO_4 at 443 K gives compound C. Treatment of C with aqueous H_2SO_4 gives compound D ($C_4H_{10}O$) which is an isomer of A. Compound D is resistant to oxidation but compound A can be easily oxidised. Identify A, B, C and D. Name the type of isomerism exhibited by A and D	3
16.	<p>Which one of the following compounds will undergo faster hydrolysis reaction by S_N1 mechanism? Justify your answer.</p> <div style="text-align: center;">  </div> <p style="text-align: center;">or $CH_3CH_2CH_2Cl$</p> <p style="text-align: center;">OR</p> <p>A compound is formed by the substitution of two chlorine atoms for two hydrogen atoms in propane. Write the structures of the isomers possible. Give the IUPAC name of the isomer which can exhibit enantiomerism.</p>	3

17.	<p>Complete the following reactions :</p> <p>(a)  + $\text{H}_2\text{N}-\text{OH} \longrightarrow$</p> <p>(b)  $\xrightarrow[\Delta]{\text{KMnO}_4, \text{H}_2\text{SO}_4}$</p> <p>(c)  + $\text{NH}_3 \xrightarrow{\text{Strong heating}}$</p>	3									
18.	<p>Give reasons for the following:</p> <p>(i) Use of aspartame as an artificial sweetener is limited to cold foods.</p> <p>(ii) Metal hydroxides are better alternatives than sodium hydrogen carbonate for treatment of acidity.</p> <p>(iii) Aspirin is used in prevention of heart attacks.</p>	3									
19.	<p>(a) Name the branched chain component of starch.</p> <p>(b) Ribose in RNA and deoxyribose in DNA differ in the structure around which carbon atom?</p> <p>(c) How many peptide linkages are present in a tripeptide?</p> <p style="text-align: center;">OR</p> <p>Give three reactions of glucose which cannot be explained by its chain structure</p>	3									
20.	<p>The following data were obtained during the first order thermal decomposition of $\text{N}_2\text{O}_5(\text{g})$ at a constant volume:</p> $2\text{N}_2\text{O}_5(\text{g}) \rightarrow 2\text{N}_2\text{O}_4(\text{g}) + \text{O}_2(\text{g})$ <table border="1" data-bbox="438 1496 1061 1668"> <thead> <tr> <th>S.No.</th><th>Time (sec.)</th><th>Total pressure(atm)</th></tr> </thead> <tbody> <tr> <td>1.</td><td>0</td><td>0.5</td></tr> <tr> <td>2.</td><td>100</td><td>0.512</td></tr> </tbody> </table> <p>Calculate the rate constant</p> <p style="text-align: center;">OR</p> <p>Two reactions of the same order have equal pre exponential factors but their activation energies differ by 24.9 kJ mol^{-1}. Calculate the ratio between the rate constants of these reactions at 27°C. (Gas constant $R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$)</p>	S.No.	Time (sec.)	Total pressure(atm)	1.	0	0.5	2.	100	0.512	3
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<p>21.</p>	<p>(a) A colloidal sol is prepared by the given method in figure. What is the charge of AgI colloidal particles in the test tube? How is the sol formed, represented?</p>  <p>(b) Explain how the phenomenon of adsorption finds application in Heterogeneous catalysis.</p> <p>(c) Which of the following electrolytes is the most effective for the coagulation of $\text{Fe}(\text{OH})_3$ sol which is a positively charged sol ? NaCl, Na_2SO_4, Na_3PO_4</p>	<p>3</p>
<p>22.</p>	<p>Describe how the following steps can be carried out?</p> <p>(a) Recovery of Gold from leached gold metal complex.. (b) Conversion of Zirconium iodide to pure Zirconium. (c) Formation of slag in the extraction of copper. (Write the chemical equations also for the reactions involved)</p> <p style="text-align: center;">OR</p> <p>Explain the use of the following:</p> <p>a) NaCN in Froth Floatation Method. b) Carbon monoxide in Mond process. c) Coke in the extraction of Zinc from Zinc Oxide</p>	<p>3</p>
<p>23.</p>	<p>Explain the following:</p> <p>(a) Out of Sc^{3+}, Co^{2+} and Cr^{3+} ions, only Sc^{3+} is colourless in aqueous solutions. (Atomic no.: $\text{Co} = 27$; $\text{Sc} = 21$ and $\text{Cr} = 24$) (b) The $E^\circ_{\text{Cu}^{2+}/\text{Cu}}$ for copper metal is positive (+0.34), unlike the remaining members of the first transition series (c) $\text{La}(\text{OH})_3$ is more basic than $\text{Lu}(\text{OH})_3$.</p>	<p>3</p>
<p>24.</p>	<p>A metal complex having composition $\text{Cr}(\text{NH}_3)_4\text{Cl}_2\text{Br}$ has been isolated in two forms A and B. The form A reacts with AgNO_3 to give a white precipitate readily soluble in dilute aqueous ammonia whereas B gives a pale yellow precipitate soluble in concentrated ammonia.</p> <p>(i) Write the formulae of isomers A and B. (ii) State the hybridisation of chromium in each of them. (iii) Calculate the magnetic moment (spin only value) of the isomer A</p>	<p>3</p>

Section-D		
25.	<p>(a) Identify A-D</p>  <p>(b) Distinguish between the following pair of compounds: (i) Aniline and Benzylamine. (ii) Methylamine and Dimethylamine.</p> <p>(c) Complete the following: $CH_3CH_2CN \xrightarrow{LiAlH_4} A \xrightarrow{0^\circ C / HNO_2} B$</p>	5
OR		
	<p>(a) Account for the following: (i) Direct nitration of aniline yields significant amount of meta derivative. (ii) Primary aromatic amines cannot be prepared by Gabriel phthalimide synthesis.</p> <p>(b) Carry out the following conversions: (i) Ethanoic acid into methanamine. (ii) Aniline to p-Bromoaniline.</p> <p>(c) Arrange the following in increasing order of basic strength: Aniline, p-nitroaniline and p-toluidine.</p>	
26.	<p>(a) A cell is prepared by dipping a zinc rod in 1M zinc sulphate solution and a silver electrode in 1M silver nitrate solution. The standard electrode potential given: $E^0_{Zn^{2+}/Zn} = -0.76 \text{ V}$, $E^0_{Ag^+/Ag} = +0.80 \text{ V}$ What is the effect of increase in concentration of Zn^{2+} on the E_{cell}?</p> <p>(b) Write the products of electrolysis of aqueous solution of NaCl with platinum electrodes.</p> <p>(c) Calculate e.m.f. of the following cell at 298 K: $Ni(s) / Ni^{2+} (0.01 \text{ M}) // Cu^{2+} (0.1 \text{ M}) / Cu (s)$ [Given $E^0_{Ni^{2+}/Ni} = -0.25 \text{ V}$, $E^0_{Cu^{2+}/Cu} = +0.34 \text{ V}$] Write the overall cell reaction.</p>	5
OR		

	<p>(a) Apply Kohlrausch law of independent migration of ions, write the expression to determine the limiting molar conductivity of calcium chloride.</p> <p>(b) Given are the conductivity and molar conductivity of NaCl solutions at 298K at different concentrations:</p> <table border="1"> <thead> <tr> <th>Concentration M</th><th>Conductivity Scm⁻¹</th><th>Molar conductivity S cm² mol⁻¹</th></tr> </thead> <tbody> <tr> <td>0.100</td><td>106.74 x 10⁻⁴</td><td>106.7</td></tr> <tr> <td>0.05</td><td>55.53 x 10⁻⁴</td><td>111.1</td></tr> <tr> <td>0.02</td><td>23.15 x 10⁻⁴</td><td>115.8</td></tr> </tbody> </table> <p>Compare the variation of conductivity and molar conductivity of NaCl solutions on dilution. Give reason.</p> <p>(c) 0.1 M KCl solution offered a resistance of 100 ohms in a conductivity cell at 298 K. If the cell constant of the cell is 1.29 cm⁻¹, calculate the molar conductivity of KCl solution.</p>	Concentration M	Conductivity Scm ⁻¹	Molar conductivity S cm ² mol ⁻¹	0.100	106.74 x 10 ⁻⁴	106.7	0.05	55.53 x 10 ⁻⁴	111.1	0.02	23.15 x 10 ⁻⁴	115.8	
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0.05	55.53 x 10 ⁻⁴	111.1												
0.02	23.15 x 10 ⁻⁴	115.8												
27.	<p>(a) Account for the following observations:</p> <ul style="list-style-type: none"> (i) SF₄ is easily hydrolysed whereas SF₆ is not easily hydrolysed (ii) Chlorine water is a powerful bleaching agent. (iii) Bi(V) is a stronger oxidising agent than Sb(V) <p>(b) What happens when</p> <ul style="list-style-type: none"> (i) White phosphorus is heated with concentrated NaOH solution in an inert atmosphere of CO₂. (ii) XeF₆ undergoes partial hydrolysis. (Give the chemical equations involved). <p style="text-align: center;">OR</p> <p>(a) What inspired N.Bartlett for carrying out reaction between Xe and PtF₆?</p> <p>(b) Arrange the following in the order of property indicated against each set:</p> <ul style="list-style-type: none"> (i) F₂, I₂, Br₂, Cl₂ (increasing bond dissociation enthalpy) (ii) NH₃, AsH₃, SbH₃, BiH₃, PH₃ (decreasing base strength) <p>(c) Complete the following equations:</p> <ul style="list-style-type: none"> (i) $Cl_2 + NaOH(\text{cold and dilute}) \rightarrow$ (ii) $Fe^{3+} + SO_2 + H_2O \rightarrow$ 	5												